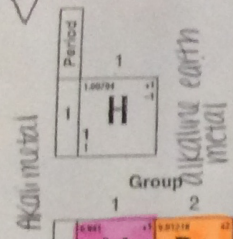


↑↑↑↑↑
Metallic Properties
↓

alkali metal
alkaline earth metal
transition metal

noble gas
halogen
metalloid

Periodic Table of the Elements



KEY

Atomic Mass → 12.011 Selected Oxidation States → -4, +2, +4

Symbol → C

Atomic Number → 6

Electron Configuration → 2-4

Relative atomic masses are based on ¹²C = 12.000

Note: Mass numbers in parentheses are mass numbers of the most stable or common isotope.

1	2	3										10	11	12	13	14	15	16	17	18	
1	2	3										10	11	12	13	14	15	16	17	18	
3	4	5										16	17	18	19	20	21	22	23	24	
11	12	13										14	15	16	17	18	19	20	21	22	
19	20	21										22	23	24	25	26	27	28	29	30	31
37	38	39										40	41	42	43	44	45	46	47	48	49
55	56	57										58	59	60	61	62	63	64	65	66	67
87	88	89										90	91	92	93	94	95	96	97	98	99

*The systematic names and symbols for elements of atomic numbers above 109 will be used until the approval of trivial names by IUPAC.

**Denotes the presence of (2-8) for elements 72 and above

1s²
alkali metals
• react quickly
• lowest density
• compounds

*radioactive
Actinides - nuclear reactors

2s²
alkaline earth metal
• react quick
• low density
• compounds

Lanthanides = magnets

58	59	60	61	62	63	64	65	66	67	68	69	70	71
Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu

Actinides

90	91	92	93	94	95	96	97	98	99	100	101	102	103
Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr

8s²
Noble gas
• stable
• non-reactive
• free elements

Transition metals
• higher MP, BP, density, strength
• unreactive
• free elements and compounds

metalloids = versatile
• metals + nonmetals
• semiconductors

9s²
Halogens
• reactive
• compounds
• salts